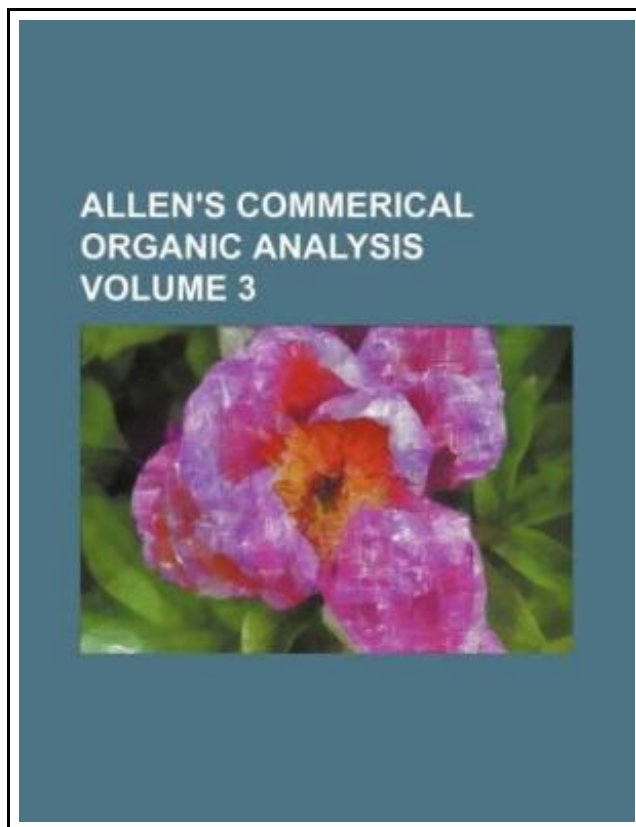


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1917 Excerpt: .phenol (even when anhydrous) containing much cresylic acid remain fluid at ordinary temperatures, and are less soluble in water and dilute alkaline liquids than is pure phenol. Cresylic acid resembles phenol in its reaction with ferric chloride. With some samples the blue colouration first produced rapidly changes to brown. Cresylic acid dissolves in dilute sodium hydroxide solution containing far less alkali than corresponds to the proportion $C_7H_6O : NaHO$, but on adding water 2 layers are formed, the lower one of which closely resembles cresylic acid, but sometimes occupies a considerably greater volume than that of the sample used. When acted on by strong nitric or sulphuric acid cresol gives similar, but not identical, products with those yielded by phenol. With excess of bromine it produces a similar body, but brominated cresol from coal-tar cresylic acid is liquid at ordinary temperatures, whereas the phenol derivative is solid.1 The following table shows the chief differences of analytical value existing between phenol and cresol. The statements were personally verified by Allen upon a sample of Calvert s No. 1 carbolic acid and a cresylic acid prepared purposely by careful fractional distillation of a sample of Calvert s No. 5 carbolic acid. The cresylic acid was further purified by dissolving it in sodium hydroxide and agitating the solution with ether to remove naphthalene and other hydrocarbons, the cresol being subsequently 1 See Allen s paper on Commercial Cresylic Acid, J. Soe. Chem. Ind., 1890, 141. recovered by addition of acid...



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